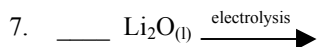
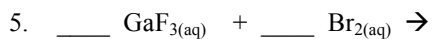


Chapter 3 problems: 64,66b,70,72,84,86

Worksheet Problems:

1. Explain how an ionic bond would be formed between potassium atoms and bromine atoms in potassium bromide.
2. Explain how a covalent bond is formed in bromine monochloride (BrCl).
3. Given the formula of iodic acid as HIO_3 , give the formula for calcium hypoiodite: _____

Predict the **products** for the following reactions, write **state symbols** for all, and **BALANCE** the equations. If **NO** Reaction is expected, briefly explain why.



Write formulas for the **reactants**, predict the **products** for the following reactions, write **state symbols** for all, and **BALANCE** the equations. If **NO** Reaction is expected, briefly explain why.

8. Liquid Octane (C_8H_{18}) is burned.

9. Potassium metal reacts with hydrogen gas (in this case, hydrogen becomes an anion).

10. An aqueous solution of iron (III) nitrate is mixed with an aqueous solution of sodium chromate.

11. Sodium phosphate reacts with cobalt (II) chloride to form a precipitate.

MOLECULAR EQUATION:

NET IONIC EQUATION: