

Isotopes 2

1. Assuming neutral atoms of the following isotopes, complete the following table.

Isotope Name	Isotope Symbol	Atomic Number (Z)	Mass Number (A)	protons	electrons	neutrons
platinum-197						
		37				50
lead-206						
	${}_{94}^{242}\text{Pu}$					
					12	14
		40	91			

2. Magnesium is composed of three naturally occurring isotopes:

Isotope	Percent Natural Abundance	Isotopic Mass (amu)
Mg – 24		23.9850
Mg – 25	10.00 %	24.9858
Mg – 26		25.9826

Calculate the % abundance of the other two isotopes.